

Historic Models

- 1 Which statement describes the earliest model of the atom?
- (1) An atom is an indivisible hard sphere.
 - (2) An atom has a small, dense nucleus.
 - (3) Electrons are negative particles in an atom.
 - (4) Electrons in an atom have wave-like properties.
- 2 The results of the gold foil experiment led to the conclusion that an atom is
- (1) mostly empty space and has a small, negatively charged nucleus
 - (2) mostly empty space and has a small, positively charged nucleus
 - (3) a hard sphere and has a large, negatively charged nucleus
 - (4) a hard sphere and has a large, positively charged nucleus
- 3 The discovery of the electron as a subatomic particle was a result of
- (1) collision theory
 - (2) kinetic molecular theory
 - (3) the gold-foil experiment
 - (4) experiments with cathode ray tubes
- 4 Given a list of atomic model descriptions:
 A: electron shells outside a central nucleus
 B: hard, indivisible sphere
 C: mostly empty space
- Which list of atomic model descriptions represents the order of historical development from the earliest to most recent?
- (1) A, B, C
 - (2) A, C, B
 - (3) B, C, A
 - (4) B, A, C
- 5 Four statements about the development of the atomic model are shown below.
- A: Electrons have wavelike properties.
 - B: Atoms have small, negatively charged particles.
 - C: The center of an atom is a small, dense nucleus.
 - D: Atoms are hard, indivisible spheres.
- Which order of statements represents the historical development of the atomic model?
- (1) C → D → A → B
 - (2) C → D → B → A
 - (3) D → B → A → C
 - (4) D → B → C → A
- 6 Which conclusion was drawn from the results of the gold foil experiment?
- (1) An atom is electrically neutral.
 - (2) An atom is mostly empty space.
 - (3) The nucleus of an atom is negatively charged.
 - (4) The electrons in an atom are located in specific shells.
- 7 As a result of the gold foil experiment, it was concluded that an atom
- (1) contains protons, neutrons, and electrons
 - (2) contains a small, dense nucleus
 - (3) has positrons and orbitals
 - (4) is a hard, indivisible sphere

Base your answers to questions 8 on the information below and on your knowledge of chemistry.

A student compares some models of the atom. These models are listed in the table below in order of development from top to bottom.

Models of the Atom		
Model	Observation	Conclusion
Dalton model	Matter is conserved during a chemical reaction.	Atoms are hard, indivisible spheres of different sizes.
Thomson model	Cathode rays are deflected by magnetic/electric fields.	Atoms have small, negatively charged particles as part of their internal structure.
Rutherford model	Most alpha particles pass straight through gold foil but a few are deflected.	An atom is mostly empty space with a small, dense, positively charged nucleus.
Bohr model	Unique spectral lines are emitted by excited gaseous elements.	Packets of energy are absorbed or emitted by atoms when an electron changes shells.

8 State one way in which the Bohr model agrees with the Thomson model.

Answer Keys

1 1

2 2

3 4

4 3

5 4

6 2

7 2

8 Allow 1 credit. Acceptable responses include, but are not limited to:

- Atoms have electrons.
- Atoms have small, negatively charged particles.
- Both models show an internal structure.
- Atoms are neutral.